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(54) **PITCH-BASED CARBON FOAM HEAT SINK WITH PHASE CHANGE MATERIAL**

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(*) Notice: Subject to any disclaimer, the term of this patent is extended or adjusted under 35 U.S.C. 154(b) by 0 days.

This patent is subject to a terminal disclaimer.

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Related U.S. Application Data

(60) Continuation of application No. 09/093,406, filed on Jun. 8, 1998, now Pat. No. 6,037,032, which is a continuation-in-part of application No. 08/921,875, filed on Sep. 2, 1997, now Pat. No. 6,033,506, and a continuation-in-part of application No. 08/923,877, filed on Sep. 2, 1997, now abandoned, application No. 09/489,799, and a continuation-in-part of application No. 09/337,027, filed on Jun. 25, 1999, now Pat. No. 6,261,485, which is a continuation of application No. 08/921,875, filed on Sep. 2, 1997, now Pat. No. 6,033,506, application No. 09/489,799, and a continuation-in-part of application No. 09/136,596, filed on Aug. 19, 1998, which is a division of application No. 08/921,875, filed on Sep. 2, 1997, now Pat. No. 6,033,506.

(51) **Int. Cl.**⁷ **B05D 7/22**; B05D 3/02; B05D 1/18

(52) **U.S. Cl.** **427/230**; 427/243; 427/244; 427/430.1

(58) **Field of Search** 427/243, 244, 427/230, 430.1

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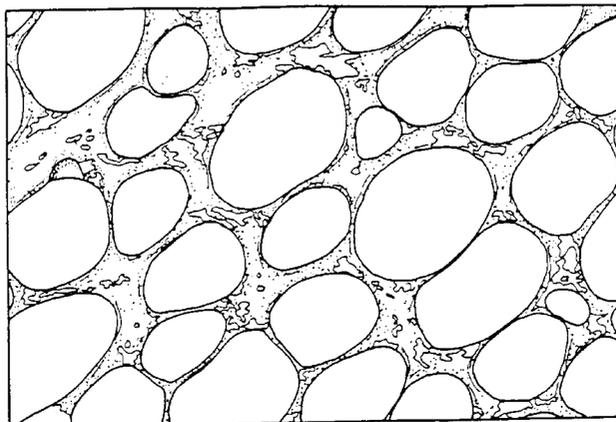
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(57) **ABSTRACT**

A process for producing a carbon foam heat sink is disclosed which obviates the need for conventional oxidative stabilization. The process employs mesophase or isotropic pitch and a simplified process using a single mold. The foam has a relatively uniform distribution of pore sizes and a highly aligned graphic structure in the struts. The foam material can be made into a composite which is useful in high temperature sandwich panels for both thermal and structural applications. The foam is encased and filled with a phase change material to provide a very efficient heat sink device.

56 Claims, 13 Drawing Sheets



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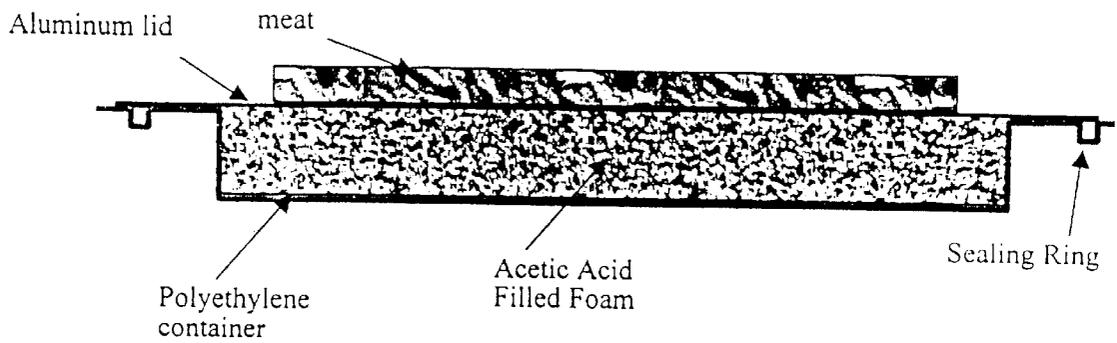


FIGURE 1

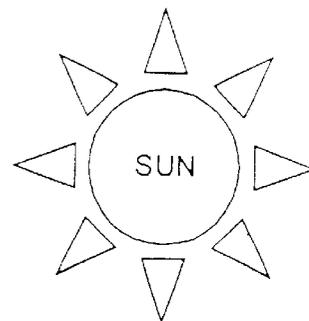
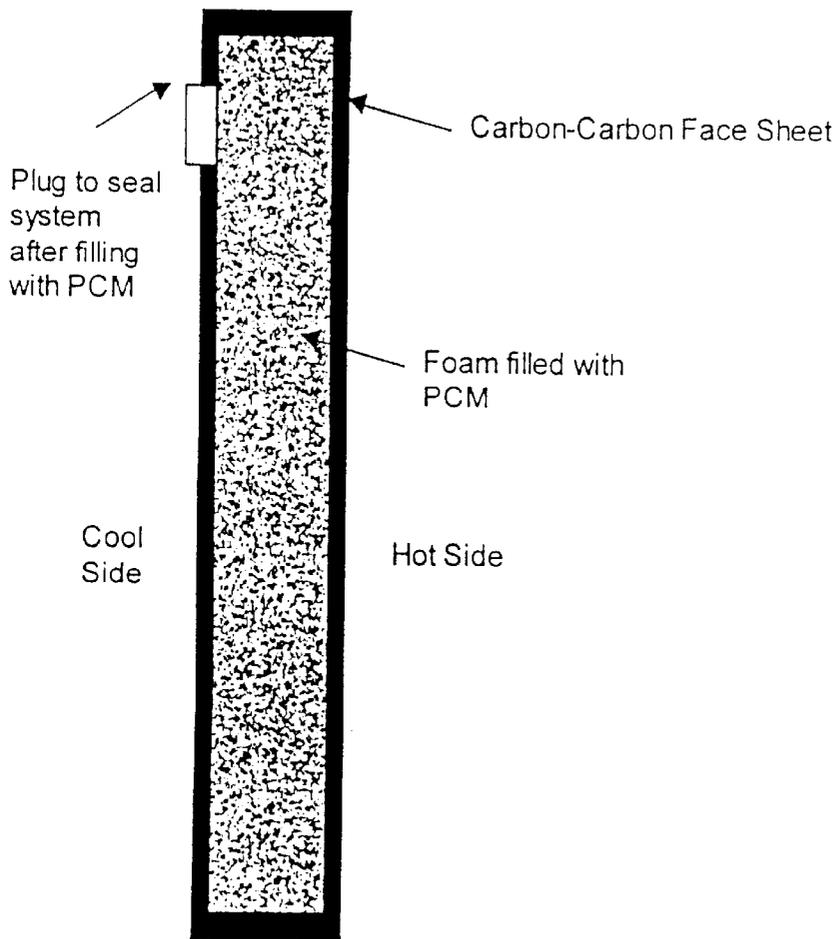
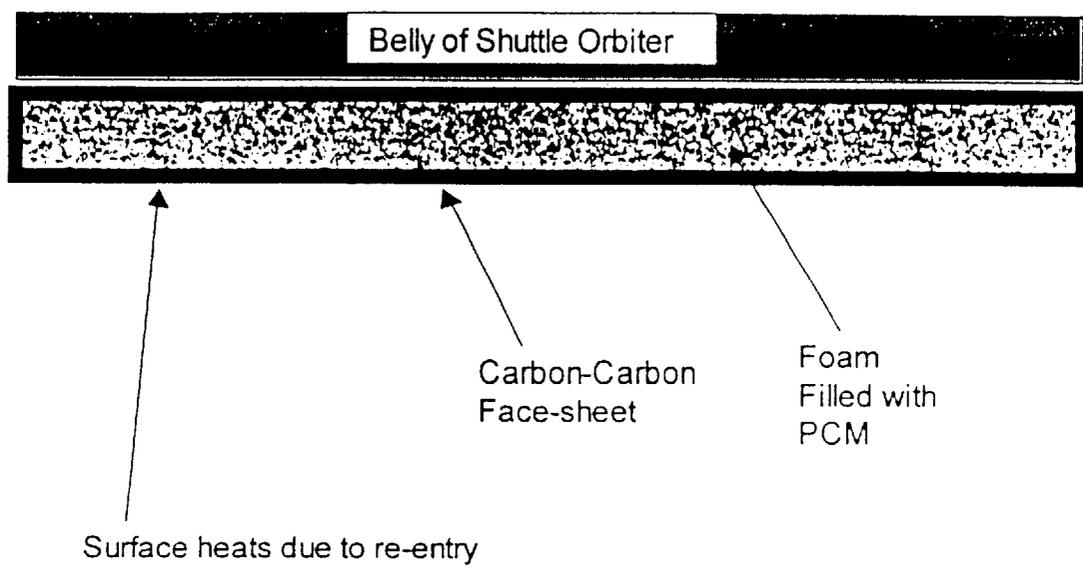


FIGURE 2



PCM keeps surface cooler than with normal materials

FIGURE 3



FIG. 4

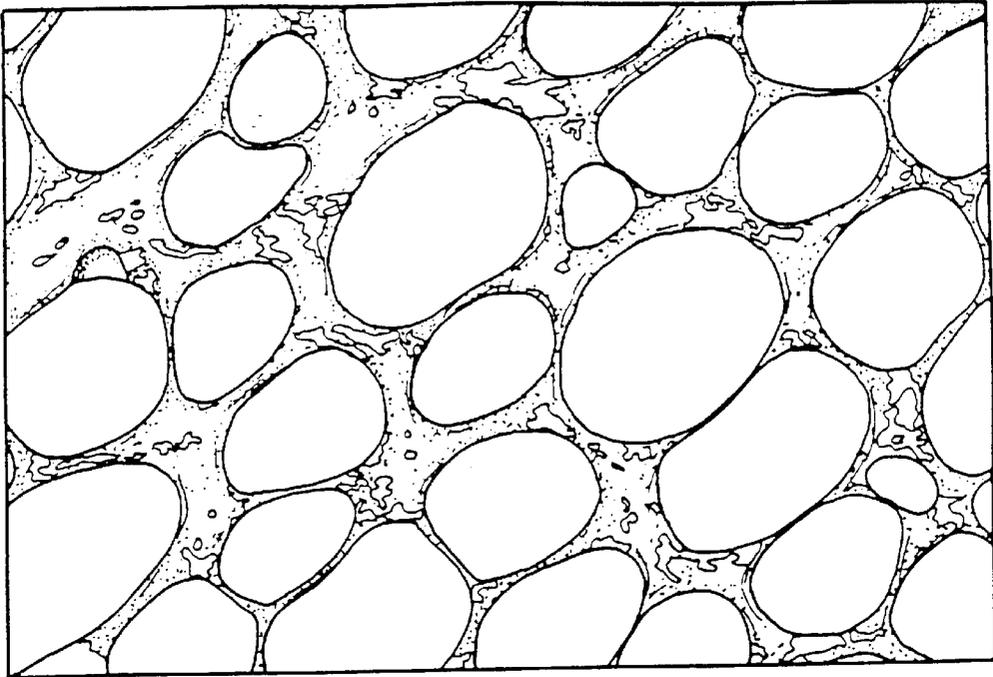


FIG. 5

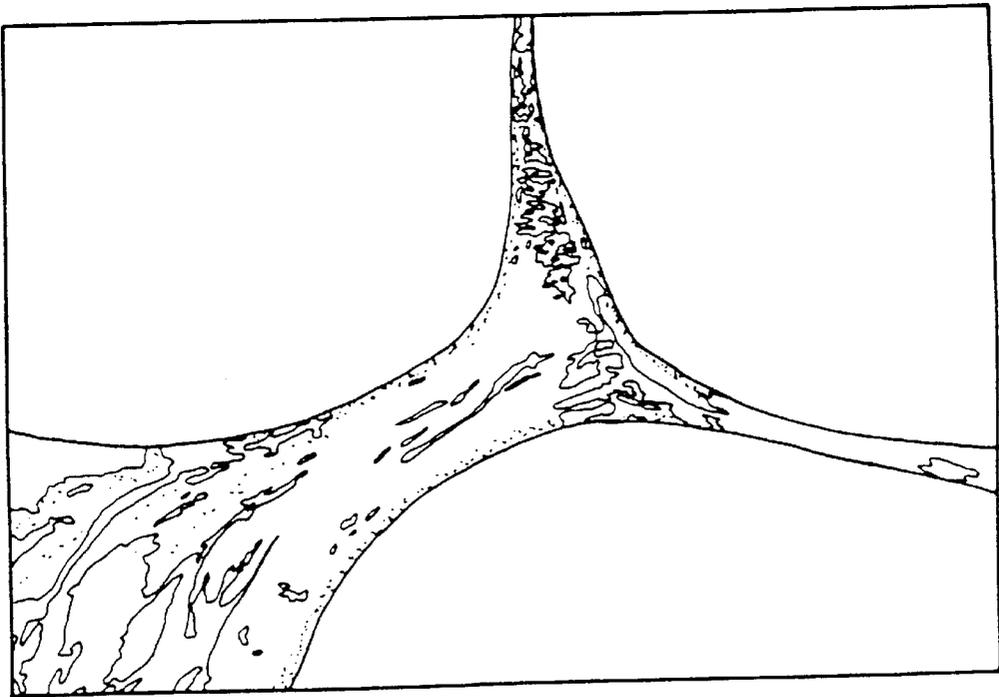


FIG. 6

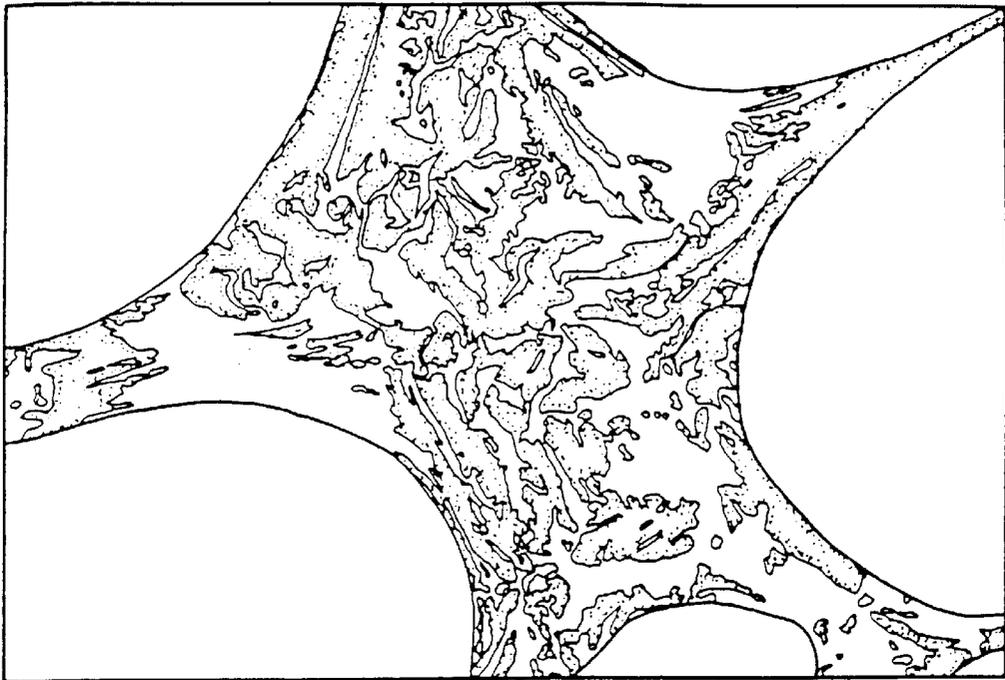


FIG. 7

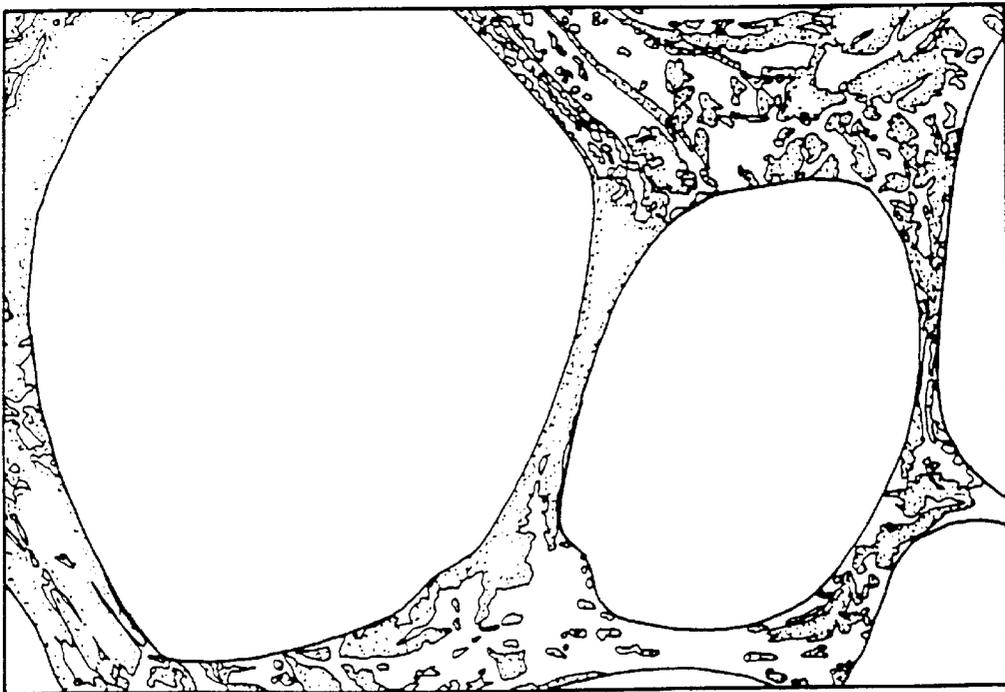


FIG. 8

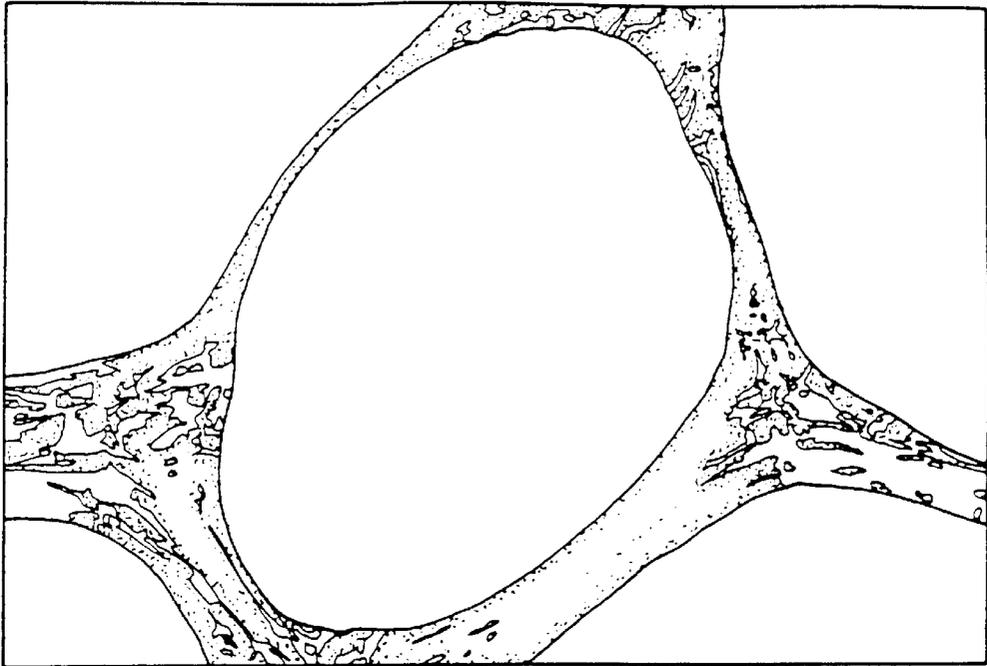


FIG. 9

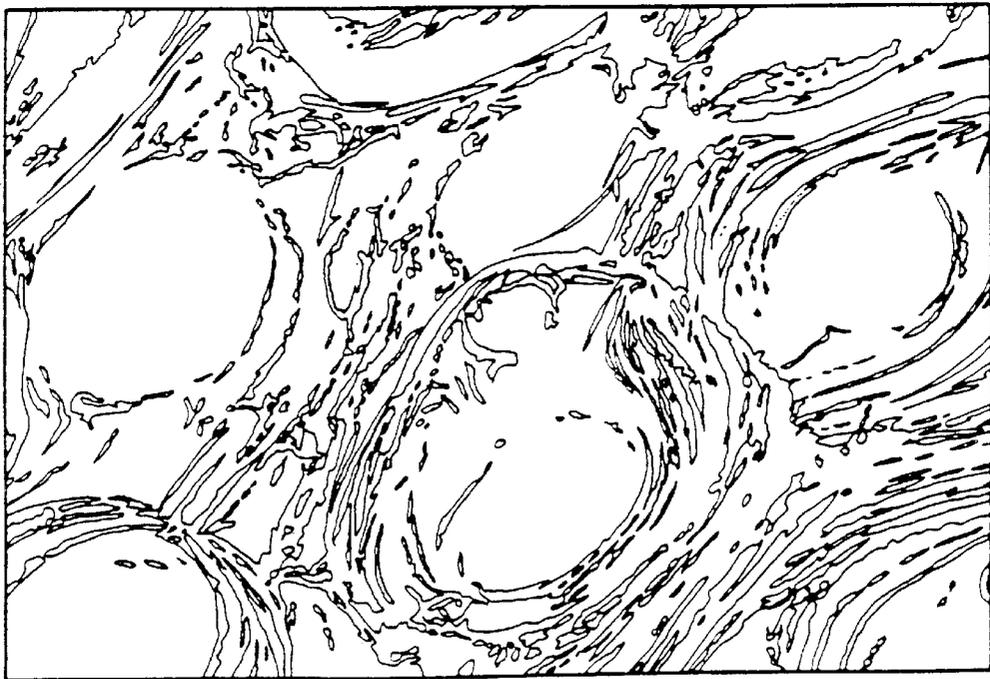


FIG. 10

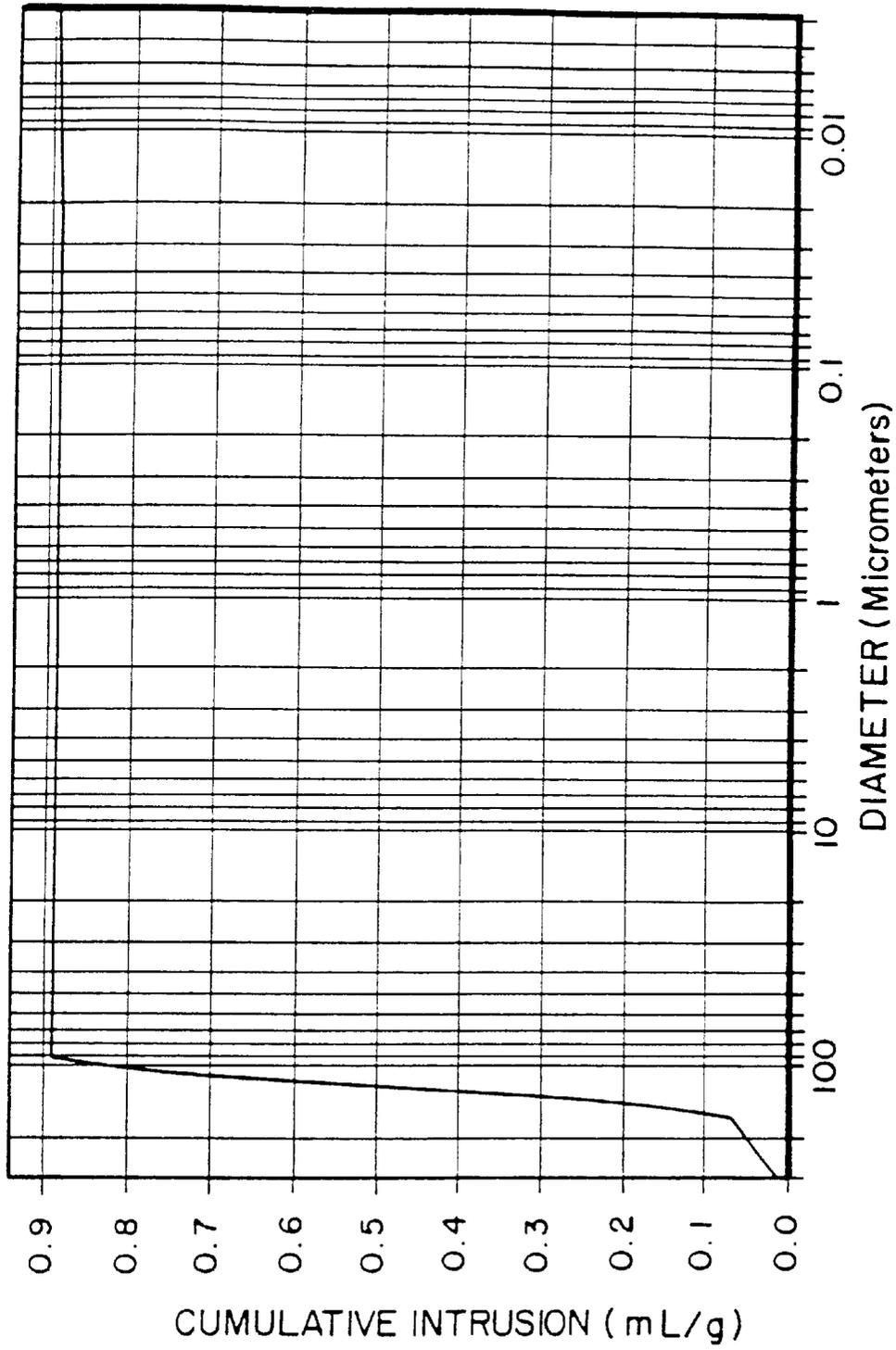


FIG. 11

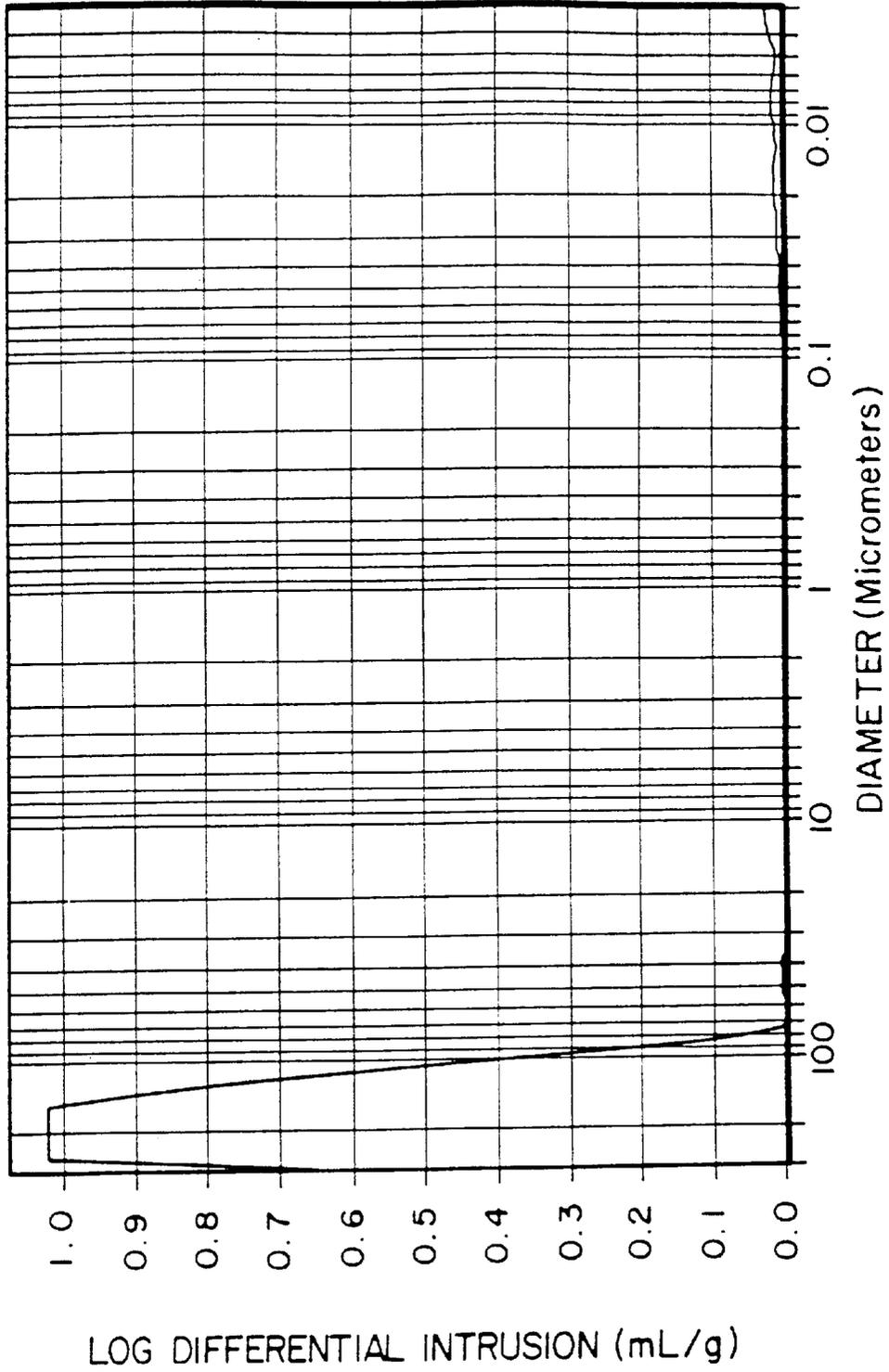


FIG. 12

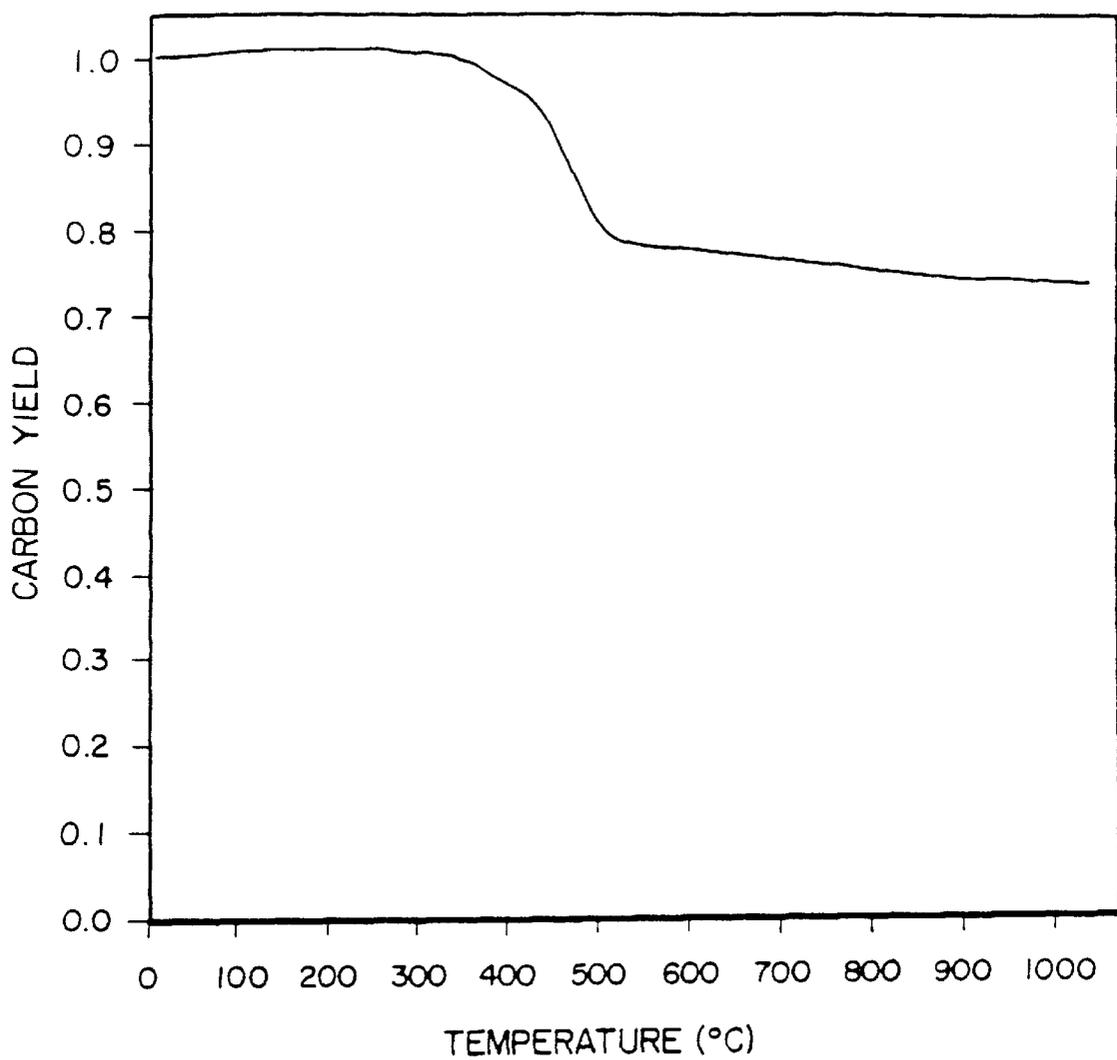


FIG. 13

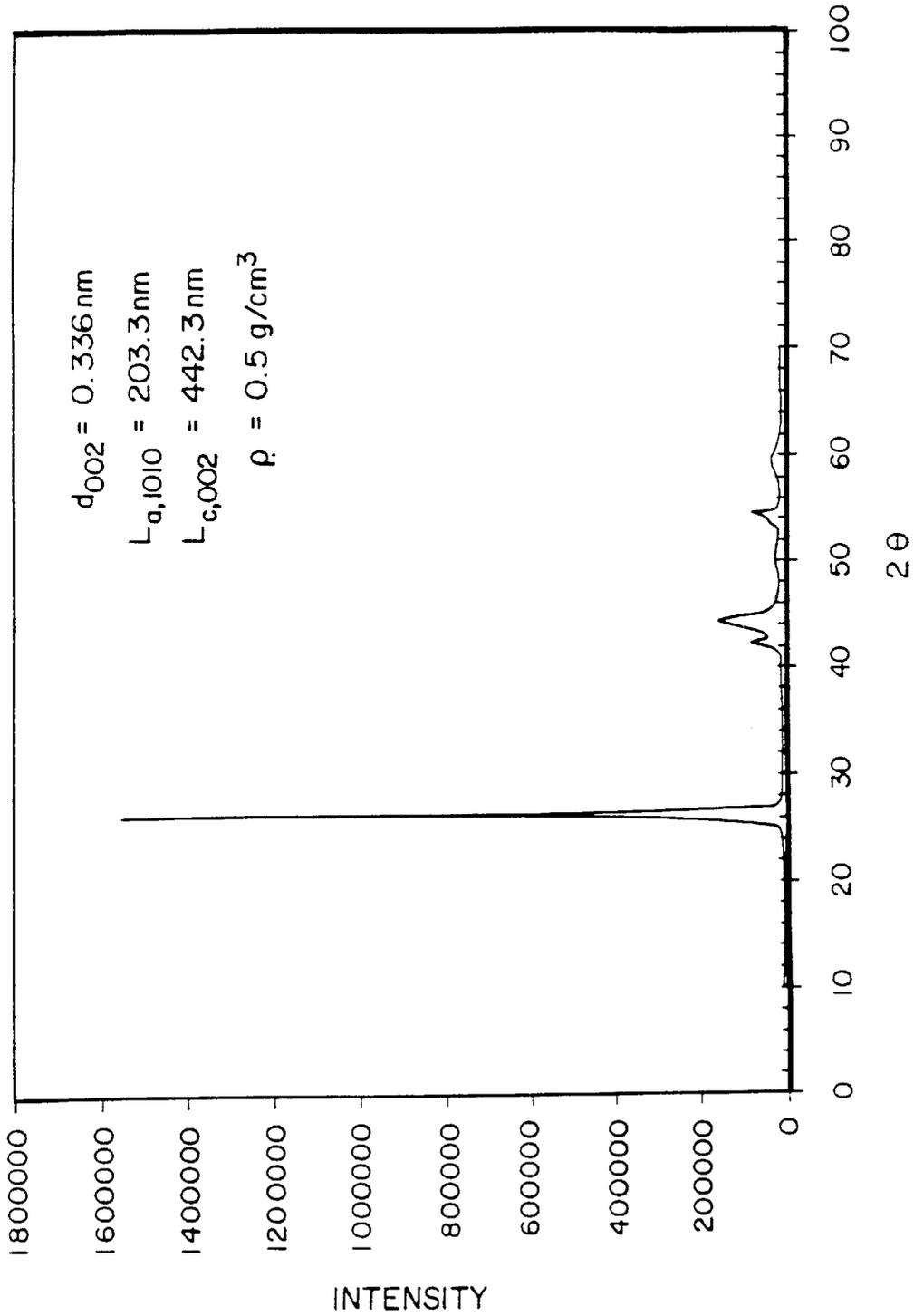
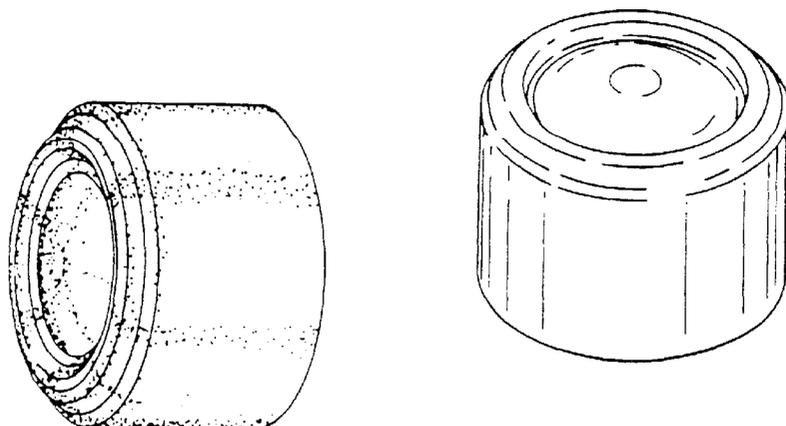
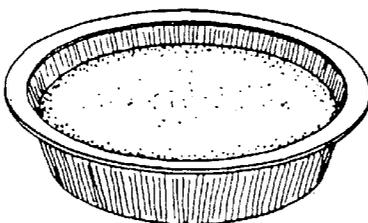


FIG. 14



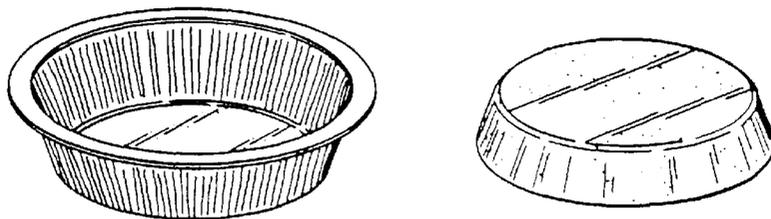
FOAM PRODUCED IN BOTTOM OF SODA CAN.

FIG. 15A



FOAM PRODUCED IN ALUMINUM WEIGHING DISH.

FIG. 15B



FOAM REMOVED FROM ALUMINUM WEIGHING DISH AND
TURNED OVER, EXPOSING SMOOTH INTEGRATED SURFACE (FACE)

FIG. 15C

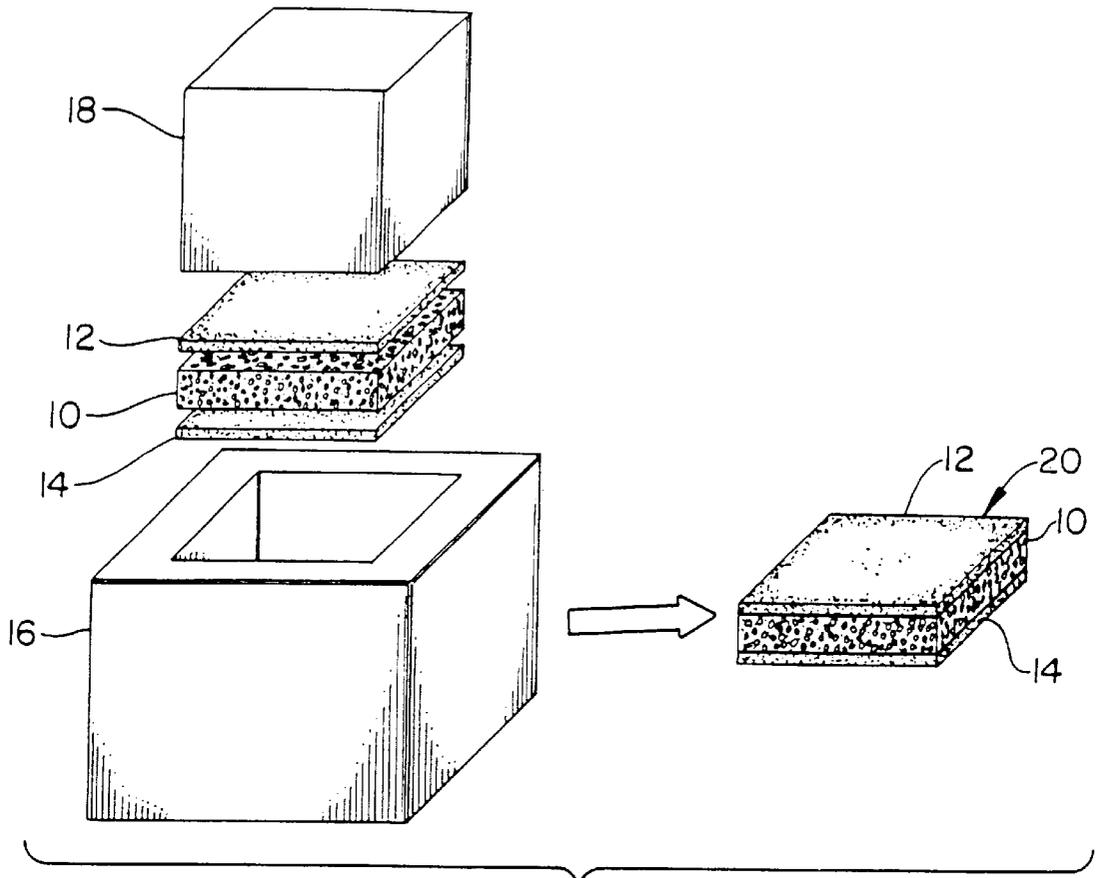


FIG. 16A

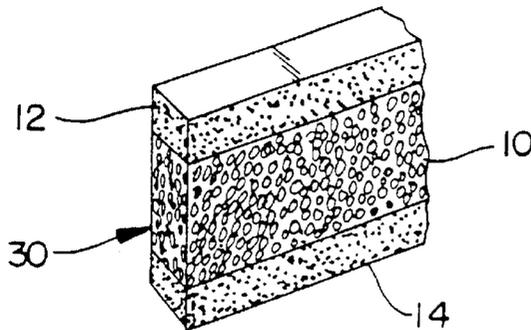


FIG. 16B

PITCH-BASED CARBON FOAM HEAT SINK WITH PHASE CHANGE MATERIAL

CROSS-REFERENCE TO RELATED APPLICATIONS

This application is (1) a continuation of U.S. patent application Ser. No. 09/093,406 filed Jun. 8, 1998 now U.S. Pat. No. 6,037,032, which itself is a continuation-in-part of an earlier filed U.S. patent application Ser. No. 08/921,875, filed on Sep. 2, 1997 now U.S. Pat. No. 6,033,506, and U.S. patent application Ser. No. 08/923,877, filed Sep. 2, 1997 now abandoned, both of which are herein incorporated in their entirety by reference and (2) a continuation-in-part of U.S. patent application No. 09/337,027 filed Jun. 25, 1999 now U.S. Pat. No. 6,261,485, which itself is a continuation of U.S. patent application Ser. No. 08/921,875 filed Sep. 2, 1997 now U.S. Pat. No. 6,033,506, and (3) a continuation-in-part of U.S. patent application Ser. No. 09/136,596 filed Aug. 19, 1998, which itself is a division of U.S. patent application Ser. No. 08/921,875 filed Sep. 2, 1997 now U.S. Pat. No. 6,033,506.

STATEMENT REGARDING FEDERALLY SPONSORED RESEARCH

The United States Government has rights in this invention pursuant to contract No. DE-AC05-96OR22464 between the United States Department of Energy and Lockheed Martin Energy Research Corporation.

BACKGROUND OF THE INVENTION

The present invention relates to porous carbon foam filled with phase change materials and encased to form a heat sink product, and more particularly to a process for producing them.

There are currently many applications that require the storage of large quantities of heat for either cooling or heating an object. Typically these applications produce heat so rapidly that normal dissipation through cooling fins, natural convection, or radiation cannot dissipate the heat quickly enough and, thus, the object over heats. To alleviate this problem, a material with a large specific heat capacity, such as a heat sink, is placed in contact with the object as it heats. During the heating process, heat is transferred to the heat sink from the hot object, and as the heat sink's temperature rises, it "stores" the heat more rapidly than can be dissipated to the environment through convection. Unfortunately, as the temperature of the heat sink rises the heat flux from the hot object decreases, due to a smaller temperature difference between the two objects. Therefore, although this method of energy storage can absorb large quantities of heat in some applications, it is not sufficient for all applications

Another method of absorbing heat is through a change of phase of the material, rather than a change in temperature. Typically, the phase transformation of a material absorbs two orders of magnitude greater thermal energy than the heat capacity of the material. For example, the vaporization of 1 gram of water at 100° C. absorbs 2,439 joules of energy, whereas changing the temperature of water from 99° C. to 100° C. only absorbs 4.21 Joules of energy. In other words, raising the temperature of 579 grams of water from 99° C. to 100° C. absorbs the same amount of heat as evaporating 1 gram of water at 100° C. The same trend is found at the melting point of the material. This phenomenon has been utilized in some applications to either absorb or evolve tremendous amounts of energy in situations where heat sinks will not work.

Although a solid block of phase change material has a very large theoretical capacity to absorb heat, the process is not a rapid one because of the difficulties of heat transfer and thus it cannot be utilized in certain applications. However, the utilization of the high thermal conductivity foam will overcome the shortcomings described above. If the high conductivity foam is filled with the phase change material, the process can become very rapid. Because of the extremely high conductivity in the struts of the foam, as heat contacts the surface of the foam, it is rapidly transmitted throughout the foam to a very large surface area of the phase change material. Thus, heat is very quickly distributed throughout the phase change material, allowing it to absorb or emit thermal energy extremely quickly without changing temperature, thus keeping the driving force for heat transfer at its maximum.

Heat sinks have been utilized in the aerospace community to absorb energy in applications such as missiles and aircraft where rapid heat generation is found. A material that has a high heat of melting is encased in a graphite or metallic case, typically aluminum, and placed in contact with the object creating the heat. Since most phase change materials have a low thermal conductivity, the rate of heat transfer through the material is limited, but this is offset by the high energy absorbing capability of the phase change. As heat is transmitted through the metallic or graphite case to the phase change material, the phase change material closest to the heat source begins to melt. Since the temperature of the phase change material does not change until all the material melts, the flux from the heat source to the phase change material remains relatively constant. However, as the heat continues to melt more phase change material, more liquid is formed. Unfortunately, the liquid has a much lower thermal conductivity, thus hampering heat flow further. In fact, the overall low thermal conductivity of the solid and liquid phase change materials limits the rate of heat absorption and, thus, reduces the efficiency of the system.

Recent developments of fiber-reinforced composites, including carbon foams, have been driven by requirements for improved strength, stiffness, creep resistance, and toughness in structural engineering materials. Carbon fibers have led to significant advancements in these properties in composites of various polymeric, metal, and ceramic matrices.

However, current applications of carbon fibers have evolved from structural reinforcement to thermal management in application ranging from high-density electronic modules to communication satellites. This has stimulated research into novel reinforcements and composite processing methods. High thermal conductivity, low weight, and low coefficient of thermal expansion are the primary concerns in thermal management applications. See Shih, Wei, "Development of Carbon-Carbon Composites for Electronic Thermal Management Applications," IDA Workshop, May 3-5, 1994, supported by A F Wright Laboratory under Contract Number F33615-93-C-2363 and A R Phillips Laboratory Contract Number F29601-93-C-0165 and Engle, G. B., "High Thermal Conductivity C/C Composites for Thermal Management," IDA Workshop, May 3-5, 1994, supported by A F Wright Laboratory under Contract F33615-93-C-2363 and A R Phillips Laboratory Contract Number F29601-93-C-0165. Such applications are striving towards a sandwich type approach in which a low density structural core material (i.e. honeycomb or foam) is sandwiched between a high thermal conductivity facesheet. Structural cores are limited to low density materials to ensure that the weight limits are not exceeded. Unfortunately, carbon foams and carbon honeycomb mate-

rials are the only available materials for use in high temperature applications (>1600° C.). High thermal conductivity carbon honeycomb materials are extremely expensive to manufacture compared to low conductivity honeycombs, therefore, a performance penalty is paid for low cost materials. High conductivity carbon foams are also more expensive to manufacture than low conductivity carbon foams, in part, due to the starting materials.

In order to produce high stiffness and high conductivity carbon foams, invariably, a pitch must be used as the precursor. This is because pitch is the only precursor which forms a highly aligned graphitic structure which is a requirement for high conductivity. Typical processes utilize a blowing technique to produce a foam of the pitch precursor in which the pitch is melted and passed from a high pressure region to a low pressure region. Thermodynamically, this produces a "Flash," thereby causing the low molecular weight compounds in the pitch to vaporize (the pitch boils), resulting in a pitch foam. See Hagar, Joseph W. and Max L. Lake, "Novel Hybrid Composites Based on Carbon Foams," *Mat. Res. Soc. Symp.*, Materials Research Society, 270:29-34 (1992); Hagar, Joseph W. and Max L. Lake, "Formulation of a Mathematical Process Model Process Model for the Foaming of a Mesophase Carbon Precursor," *Mat. Res. Soc. Symp.*, Materials Research Society, 270:35-40 (1992); Gibson, L. J. and M. F. Ashby, *Cellular Solids: Structures & Properties*, Pergamon Press, New York (1988); Gibson, U., *Mat. Sci. and Eng A110*, 1(1989); Knippenberg and B. Lersmacher, *Phillips Tech. Rev.*, 36(4), (1976); and Bonzom, A., P. Crepau and E. J. Moutard, U.S. Pat. No. 4,276,246, (1981). Then, the pitch foam must be oxidatively stabilized by heating in air (or oxygen) for many hours, thereby, crosslinking the structure and "stabilizing" the pitch so it does not melt during carbonization. See Hagar, Joseph W. and Max L. Lake, "Formulation of a Mathematical Process Model Process Model for the Foaming of a Mesophase Carbon Precursor," *Mat. Res. Soc. Symp.*, Materials Research Society, 270:35-40 (1992); and White, J. L., and P. M. Shaeffer, *Carbon*, 27:697 (1989). This is a time consuming step and can be an expensive step depending on the part size and equipment required. The "stabilized" or oxidized pitch is then carbonized in an inert atmosphere to temperatures as high as 1100° C. Then, graphitization is performed at temperatures as high as 3000° C. to produce a high thermal conductivity graphitic structure, resulting in a stiff and very thermally conductive foam.

Other techniques utilize a polymeric precursor, such as phenolic, urethane, or blends of these with pitch. See Hagar, Joseph W. and Max L. Lake, "Idealized Strut Geometries for Open-Cell Foams," *Mat. Res. Soc. Symp.*, Materials Research Society, 270:41-46 (1992); Aubert, J. W., *MRS Symposium Proceedings*, 207:117-127 (1990); Cowland, F. C. and J. C. Lewis, *J. of Mat. Sci.*, 2:507-512 (1967); and Noda, T., Inagaki and S. Yamada, *J. of Non-Crystalline Solids*, 1:285-302, (1969). High pressure is applied and the sample is heated. At a specified temperature, the pressure is released, thus causing the liquid to foam as volatile compounds are released. The polymeric precursors are cured and then carbonized without a stabilization step. However, these precursors produce a "glassy" or vitreous carbon which does not exhibit graphitic structure and, thus, has low thermal conductivity and low stiffness. See Hagar, Joseph W. and Max L. Lake, "Idealized Strut Geometries for Open-Cell Foams," *Mat. Res. Soc. Symp.*, Materials Research Society, 270:41-46 (1992).

In either case, once the foam is formed, it is then bonded in a separate step to the facesheet used in the composite. This can be an expensive step in the utilization of the foam.

The extraordinary mechanical properties of commercial carbon fibers are due to the unique graphitic morphology of the extruded filaments. See Edie, D. D., "Pitch and Mesophase Fibers," in *Carbon Fibers, Filaments and Composites*, Figueiredo (editor), Kluwer Academic Publishers, Boston, pp. 43-72 (1990). Contemporary advanced structural composites exploit these properties by creating a disconnected network of graphitic filaments held together by an appropriate matrix. Carbon foam derived from a pitch precursor can be considered to be an interconnected network of graphitic ligaments or struts, as shown in FIG. 4. As such interconnected networks, they represent a potential alternative as reinforcement in structural composite materials.

The process of this invention overcomes current manufacturing limitations by avoiding a "blowing" or "pressure release" technique to produce the foam. Furthermore, an oxidation stabilization step is not required, as in other methods used to produce pitch based carbon foams with a highly aligned graphitic structure. This process is less time consuming, and therefore, will be lower in cost and easier to fabricate. The foam can be produced with an integrated sheet of high thermal conductivity carbon on the surface of the foam, thereby producing a carbon foam with a smooth sheet on the surface to improve heat transfer.

SUMMARY OF THE INVENTION

An object of the present invention is production of encased high thermal conductivity porous carbon foam filled with a phase change material wherein tremendous amounts of thermal energy are stored and emitted very rapidly. The porous foam is filled with a phase change material (PCM) at a temperature close to the operating temperature of the device. As heat is added to the surface, from a heat source such as a computer chip, friction due to re-entry through the atmosphere, or radiation such as sunlight, it is transmitted rapidly and uniformly throughout the foam and then to the phase change material. As the material changes phase, it absorbs orders of magnitude more energy than non-PCM material due to transfer of the latent heat of fusion or vaporization. Conversely, the filled foam can be utilized to emit energy rapidly when placed in contact with a cold object.

Non-limiting embodiments disclosed herein are a device to rapidly thaw frozen foods or freeze thawed foods, a design to prevent overheating of satellites or store thermal energy as they experience cyclic heating during orbit, and a design to cool leading edges during hypersonic flight or re-entry from space.

Another object of the present invention is to provide carbon foam and a composite from a mesophase or isotropic pitch such as synthetic, petroleum or coal-tar based pitch. Another object is to provide a carbon foam and a composite from pitch which does not require an oxidative stabilization step.

These and other objectives are accomplished by a method of producing carbon foam heat sink wherein an appropriate mold shape is selected and preferably an appropriate mold release agent is applied to walls of the mold. Pitch is introduced to an appropriate level in the mold, and the mold is purged of air by applying a vacuum, for example. Alternatively, an inert fluid could be employed. The pitch is heated to a temperature sufficient to coalesce the pitch into a liquid which preferably is of about 50° C. to about 100° C. above the softening point of the pitch. The vacuum is released and an inert fluid applied at a static pressure up to

about 1000 psi. The pitch is heated to a temperature sufficient to cause gases to evolve and foam the pitch. The pitch is further heated to a temperature sufficient to coke the pitch and the pitch is cooled to room temperature with a simultaneous and gradual release of pressure. The foam is then filled with a phase change material and encased to produce an efficient heat storage product.

In another aspect, the previously described steps are employed in a mold composed of a material such that the molten pitch does not adhere to the surface of the mold.

In yet another aspect, the objectives are accomplished by the carbon foam product produced by the methods disclosed herein including a foam product with a smooth integral facesheet.

In still another aspect a carbon foam composite product is produced by adhering facesheets to the carbon foam produced by the process of this invention.

FIG. 1 is section cut of a heat sink device for thawing food using acetic acid as the phase change material.

FIG. 2 is a section cut of a heat sink to prevent overheating of satellites during cyclic orbits.

FIG. 3 is a section cut of a heat sink used on the leading edge of a shuttle orbiter.

FIG. 4 is a micrograph illustrating typical carbon foam with interconnected carbon ligaments and open porosity.

FIGS. 5-9 are micrographs of pitch-derived carbon foam graphitized at 2500° C. and at various magnifications.

FIG. 10 is a SEM micrograph of carbon foam produced by the process of this invention.

FIG. 11 is a chart illustrating cumulative intrusion volume versus pore diameter.

FIG. 12 is a chart illustrating log differential intrusion volume versus pore diameter.

FIG. 13 is a graph illustrating the temperatures at which volatiles are given off from raw pitch.

FIG. 14 is an X-ray analysis of the graphitized foam produced by the process of this invention.

FIGS. 15A-C are photographs illustrating foam produced with aluminum crucibles and the smooth structure or face sheet that develops.

FIG. 16A is a schematic view illustrating the production of a carbon foam composite made in accordance with this invention.

FIG. 16B is a perspective view of the carbon foam composite of this invention.

DETAILED DESCRIPTION OF THE INVENTION

In order to illustrate the carbon foam heat sink product of this invention, the following examples are set forth. They are not intended to limit the invention in any way.

Example 1: Device for Thawing Food

Acetic acid has a heat of melting of 45 J/g at a melting point of 11° C. The heat of melting of food, primarily ice, is roughly 79 J/g at 0° C. Therefore, take a block of foam and fill it with liquid acetic acid at room temp. The foam will be encased in a box made from an insulating polymer such as polyethylene on all sides except the top. The top of the foam/acetic acid block will be capped with a high thermal conductivity aluminum plate that snaps into place thus sealing the foam/acetic acid inside the polymer case (illustrated in FIG. 1). If the foam block is 10-in.×15-in.×

0.5-in. thick, the mass of foam is 614 grams. The mass of acetic acid that fills the foam is roughly 921 grams. Therefore, when a piece of frozen meat is placed in contact with the top of the aluminum block, the foam will cool to the freezing point of the acetic acid (11° C.). At this point, the heat given off from the acetic acid as it freezes (it also remains at 11° C.) will be equivalent to 49 KJ. This heat is rapidly transferred to the frozen meat as it thaws (it also remains at 0° C.). This amount of heat is sufficient to melt roughly 500 grams (1 lb.) of meat.

Example 2: Heat Sink To Prevent Overheating Of Satellites During Cyclic Orbits

Produce a carbon-carbon composite with the foam in which the foam is a core material with carbon-carbon face sheets (FIG. 2). Fill the foam core with a suitable phase change material, such as a paraffin wax, that melts around the maximum operating temperature of the satellite components. One method to perform this would be to drill a hole in one surface of the carbon-carbon face sheets and vacuum fill the phase change material in the liquid state into the porous foam. Once filled, the sample can be cooled (the phase change material solidifies) and the hole can be plugged with an epoxy or screw-type cap. The epoxy and any other sealant must be able to withstand the operating temperature of the application. The foam-core composite will then be mounted on the side of the satellite that is exposed to the sun during orbit. As the satellite orbits the earth and is exposed to the sun, the radiant energy from the sun will begin to heat the composite panel to the melting point of the phase change material. At this point, the panel will not increase in temperature as the phase change material melts. The amount of radiant energy the panel can absorb will be dependent on the thickness and outer dimensions of the panel. This can be easily calculated and designed through knowledge of the orbit times of the satellite such that the material never completely melt and, thus, never exceeds the melt temperature. Then, when the satellite leaves the view of the sun, it will begin to radiate heat to space and the phase change material will begin to freeze. The cycle will repeat itself once the satellite comes into view of the sun once again.

Example 3: Heat Sink for Leading Edges

Currently, the shuttle orbiter experiences extreme heats during reentry. Specifically, the leading edges of the craft can reach 1800° C. and the belly of the craft can reach temperatures as high as 1200° C. If a foam core composite panel is placed at the surface of the leading edges and at the surface of the belly (FIG. 3), it would be able to absorb enough energy to dramatically reduce the maximum temperature of the hot areas. This also would permit a faster re-entry or (steeper glide slope) and maintain the current maximum temperatures. In this case the phase change material would most likely be an alloy, e.g. germanium-silicon, which melts around 800-900° C. and does not vaporize until much higher than the maximum temperature of the craft.

For example, Germanium has a heat of formation (heat of melting) of 488 J/g. This would require 1.0 Kg of Germanium to reduce the temperature of 1 Kg of existing carbon/carbon heat-shield by 668° C. In other words, if the existing carbon-carbon were replaced pound-for-pound with germanium filled foam, the maximum temperature of the heat shield would only be about 1131° C. instead of about 1800° C. during reentry, depending on the duration of thermal loading.

Example 4

Pitch powder, granules, or pellets are placed in a mold with the desired final shape of the foam. These pitch materials can be solvated if desired. In this example, Mitsubishi ARA-24 mesophase pitch was utilized. A proper mold release agent or film is applied to the sides of the mold to allow removal of the part. In this case, Boron Nitride spray and Dry Graphite Lubricant were separately used as a mold release agent. If the mold is made from pure aluminum, no mold release agent is necessary since the molten pitch does not adhere to the aluminum and, thus, will not stick to the mold. Similar mold materials may be found that the pitch does not adhere and, thus, they will not need mold release. The sample is evacuated to less than 1 torr and then heated to a temperature approximately 50 to 100° C. above the softening point. In this case where Mitsubishi ARA24 mesophase pitch was used, 300° C. was sufficient. At this point, the vacuum is released to a nitrogen blanket and then a pressure of up to 1000 psi is applied. The temperature of the system is then raised to 800° C., or a temperature sufficient to coke the pitch which is 500° C. to 1000° C. This is performed at a rate of no greater than 5° C./min. and preferably at about 2° C./min. The temperature is held for at least 15 minutes to achieve an assured soak and then the furnace power is turned off and cooled to room temperature. Preferably the foam was cooled at a rate of approximately 1.5° C./min. with release of pressure at a rate of approximately 2 psi/min. Final foam temperatures for three product runs were 500° C., 630° C. and 800° C. During the cooling cycle, pressure is released gradually to atmospheric conditions. The foam was then heat treated to 1050° C. (carbonized) under a nitrogen blanket and then heat treated in separate runs to 2500° C. and 2800° C. (graphitized) in Argon.

Carbon foam produced with this technique was examined with photomicrography, scanning electron microscopy (SEM), X-ray analysis, and mercury porosimetry. As can be seen in the FIGS. 5–10, the isochromatic regions under cross-polarized light indicate that the struts of the foam are completely graphitic. That is, all of the pitch was converted to graphite and aligned along the axis of the struts. These struts are also similar in size and are interconnected throughout the foam. This would indicate that the foam would have high stiffness and good strength. As seen in FIG. 10 by the SEM micrograph of the foam, the foam is open cellular meaning that the porosity is not closed. FIGS. 11 and 12 are results of the mercury porosimetry tests. These tests indicate that the pore sizes are in the range of 90–200 microns.

A thermogravimetric study of the raw pitch was performed to determine the temperature at which the volatiles are evolved. As can be seen in FIG. 14, the pitch loses nearly 20% of its mass fairly rapidly in the temperature range between about 420° C. and about 480° C. Although this was performed at atmospheric pressure, the addition of 1000 psi pressure will not shift this effect significantly. Therefore, while the pressure is at 1000 psi, gases rapidly evolved during heating through the temperature range of 420° C. to 480° C. The gases produce a foaming effect (like boiling) on the molten pitch. As the temperature is increased further to temperatures ranging from 500° C. to 1000° C. (depending on the specific pitch), the foamed pitch becomes coked (or rigid), thus producing a solid foam derived from pitch. Hence, the foaming has occurred before the release of pressure and, therefore, this process is very different from previous art.

Samples from the foam were machined into specimens for measuring the thermal conductivity. The bulk thermal con-

ductivity ranged from 58 W/m•K to 106 W/m•K. The average density of the samples was 0.53 g/cm³. When weight is taken into account, the specific thermal conductivity of the pitch derived from foam is over 4 times greater than that of copper. Further derivations can be utilized to estimate the thermal conductivity of the struts themselves to be nearly 700 W/m•K. This is comparable to high thermal conductivity carbon fibers produced from this same ARA24 mesophase pitch.

X-ray analysis of the foam was performed to determine the crystalline structure of the material. The x-ray results are shown in FIG. 14. From this data, the graphene layer spacing (d_{002}) was determined to be 0.336 nm. The coherence length ($L_{g,100}$) was determined to be 203.3 nm and the stacking height was determined to be 442.3 nm.

The compression strength of the samples were measured to be 3.4 MPa and the compression modulus was measured to be 73.4 MPa. The foam sample was easily machined and could be handled readily without fear of damage, indicating good strength.

It is important to note that when this pitch is heated in a similar manner, but only under atmospheric pressure, the pitch foams dramatically more than when under pressure. In fact, the resulting foam is so fragile that it could not even be handled to perform tests. Molding under pressure serves to limit the growth of the cells and produces a usable material.

Example 5

An alternative to the method of Example 4 is to utilize a mold made from aluminum. In this case two molds were used, an aluminum weighing dish and a sectioned soda can. The same process as set forth in Example 4 is employed except that the final coking temperature was only 630° C., so as to prevent the aluminum from melting.

FIGS. 15 A–C illustrate the ability to utilize complex shaped molds for producing complex shaped foam. In one case, shown in FIG. 15A, the top of a soda can was removed and the remaining can used as a mold. No release agent was utilized. Note that the shape of the resulting part conforms to the shape of the soda can, even after graphitization to 2800° C. This demonstrates the dimensional stability of the foam and the ability to produce near net shaped parts.

In the second case, as shown in FIGS. 15B and C employing an aluminum weight dish, a very smooth surface was formed on the surface contacting the aluminum. This is directly attributable to the fact that the molten pitch does not adhere to the surface of the aluminum. This would allow one to produce complex shaped parts with smooth surfaces so as to improve contact area for bonding or improving heat transfer. This smooth surface will act as a face sheet and, thus, a foam-core composite can be fabricated in-situ with the fabrication of the face sheet. Since it is fabricated together and an integral material no interface joints result, thermal stresses will be less, resulting in a stronger material.

The following examples illustrate the production of a composite material employing the foam of this invention.

Example 6

Pitch derived carbon foam was produced with the method described in Example 4. Referring to FIG. 16A the carbon foam 10 was then machined into a block 2"×2"×½". Two pieces 12 and 14 of a prepreg comprised of Hercules AS4 carbon fibers and ICI Fibirite Polyetheretherkeytone thermoplastic resin also of 2"×2"×½" size were placed on the top and bottom of the foam sample, and all was placed in a

matched graphite mold 16 for compression by graphite plunger 18. The composite sample was heated under an applied pressure of 100 psi to a temperature of 380° C. at a rate of 5° C./min. The composite was then heated under a pressure of 100 psi to a temperature of 650° C. The foam core sandwich panel generally 20 was then removed from the mold and carbonized under nitrogen to 1050° C. and then graphitized to 2800° C., resulting in a foam with carbon-carbon facesheets bonded to the surface. The composite generally 30 is shown in FIG. 16B.

Example 7

Pitch derived carbon foam was produced with the method described in Example 4. It was then machined into a block 2"×2"×½". Two pieces of carbon-carbon material, 2"×2"×½", were coated lightly with a mixture of 50% ethanol, 50% phenolic Durez© Resin available from Occidental Chemical Co. The foam block and carbon-carbon material were positioned together and placed in a mold as indicated in Example 6. The sample was heated to a temperature of 150° C. at a rate of 5° C./min and soaked at temperature for 14 hours. The sample was then carbonized under nitrogen to 1050° C. and then graphitized to 2800° C., resulting in a foam with carbon-carbon facesheets bonded to the surface. This is also shown generally at 30 in FIG. 16B.

Example 8

Pitch derived carbon foam was produced with the method described in Example 4. The foam sample was then densified with carbon by the method of chemical vapor infiltration for 100 hours. The density increased to 1.4 g/cm³, the flexural strength was 19.5 MPa and the flexural modulus was 2300 MPa. The thermal conductivity of the raw foam was 58 W/m•K and the thermal conductivity of the densified foam was 94 W/m•K.

Example 9

Pitch derived carbon foam was produced with the method described in Example 4. The foam sample was then densified with epoxy by the method of vacuum impregnation. The epoxy was cured at 150° C. for 5 hours. The density increased to 1.37 g/cm³ and the flexural strength was measured to be 19.3 MPa.

Other possible embodiments may include materials, such as metals, ceramics, plastics, or fiber-reinforced plastics bonded to the surface of the foam of this invention to produce a foam core composite material with acceptable properties. Additional possible embodiments include ceramics, glass, or other materials impregnated into the foam for densification.

Based on the data taken to date from the carbon foam material, several observations can be made outlining important features of the invention that include:

1. Pitch-based carbon foam can be produced without an oxidative stabilization step, thus saving time and costs.
2. High graphitic alignment in the struts of the foam is achieved upon graphitization to 2500° C., and thus high thermal conductivity and stiffness will be exhibited by the foam, making them suitable as a core material for thermal applications.
3. High compressive strengths should be achieved with mesophase pitch-based carbon foams, making them suitable as a core material for structural applications.
4. Foam core composites can be fabricated at the same time as the foam is generated, thus saving time and costs.

5. Rigid monolithic preforms can be made with significant open porosity suitable for densification by the Chemical Vapor Infiltration method of ceramic and carbon infiltrants.

6. Rigid monolithic preforms can be made with significant open porosity suitable for activation, producing a monolithic activated carbon.

7. It is obvious that by varying the pressure applied, the size of the bubbles formed during the foaming will change and, thus, the density, strength, and other properties can be affected,

The following alternative procedures and products can also be effected by the process of this invention:

1. Fabrication of preforms with complex shapes for densification by CVI or Melt Impregnation.
2. Activated carbon monoliths with high thermal conductivity.
3. Optical absorbent.
4. Low density heating elements.
5. Firewall Material
6. Low secondary electron emission targets for high-energy physics applications.

The present invention provides for the manufacture of pitch-based carbon foam heat sink for structural and thermal composites. The process involves the fabrication of a graphitic foam from a mesophase or isotropic pitch which can be synthetic, petroleum, or coal-tar based. A blend of these pitches can also be employed. The simplified process utilizes a high pressure high temperature furnace and thereby, does not require an oxidative stabilization step. The foam has a relatively uniform distribution of pore sizes (~100 microns), very little closed porosity, and density of approximately 0.53 g/cm³. The mesophase pitch is stretched along the struts of the foam structure and thereby produces a highly aligned graphitic structure in the struts. These struts will exhibit thermal conductivities and stiffness similar to the very expensive high performance carbon fibers (such as P-120 and K1100). Thus, the foam will exhibit high stiffness and thermal conductivity at a very low density (~0.5 g/cc). This foam can be formed in place as a core material for high temperature sandwich panels for both thermal and structural applications, thus reducing fabrication time. By utilizing an isotropic pitch, the resulting foam can be easily activated to produce a high surface area activated carbon. The activated carbon foam will not experience the problems associated with granules such as attrition, channeling, and large pressure drops.

The following Examples demonstrate the evaporative cooling effect on the previously described carbon foam when contacted with different working fluids as represented by acetone, ethanol and water. These Examples are not intended to limit the invention in any way. The foamed carbon was doused or partially immersed in the working fluid. Upon removal from the working fluid, and as indicated in Examples 10–16, the foam sample was placed in a vacuum furnace with a thermocouple penetrating the foam sample. The foam temperature was monitored as a function of time and pressure (vacuum). The ambient laboratory temperature was approximately 21° C.

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Example 10: Acetone

Time (minutes)	Pressure (Torr)	Temperature (° C.)
0	740	13.5
1	29	-37.5
2	29	-46.7
3	1.2	-51.8
4	1.2	-53.4

When the sample was removed from the vacuum furnace it was noted that ice had formed, presumably from moisture condensed from the furnace atmosphere, or desorbed from the foam.

Example 11: Ethanol

Time (minutes)	Pressure (Torr)	Temperature (° C.)
0	740	20.5
1	29	5.3
2	29	-14.7
3	1.2	-21.7
4	1.2	-25.1
5	1.1	-26.8
6	1.0	-28.6

Example 12: Water

Time (minutes)	Pressure (Torr)	Temperature (° C.)
0	740	20.5
1	29	16.4
2	29	16.5
3	29	16.6
4	29	14.6
5	29	12.9
6	29	10.5
7	29	2.6
8	29	-1.5
9	29	-5.5

In the instance of Example 12, the sample was immersed in water in vacuum to ensure that the foam was saturated. This probably allowed an excess of water to penetrate the sample and reduced the exposed foam surface area available for evaporation. Moreover, the resultant high water partial pressure in the furnace made it impossible to attain good vacuum in a reasonable time. Consequently, the experiment was repeated in Example 13, but with substantially less water applied to the foam.

Example 13: Water (Repeat)

Time (minutes)	Pressure (Torr)	Temperature (° C.)
0	740	19.9
1	29	14.5
2	29	0.3
3	29	-5.5

In this case, sub-zero temperatures were attained in a much shorter time than for Example 12.

A further series of tests as set forth in Examples 14-16 were performed to show the effect of evaporative cooling at

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atmospheric pressure and temperature. The foamed carbon sample was placed in a petri dish. A thermocouple was located in a hole machined into the foam. The carbon foam was doused with the working fluid until the bottom of the petri dish was completely covered with the working fluid. The resultant foam temperature was then noted as a function of time.

Example 14: Acetone

Time (minutes)	Temperature (° C.)
0	21.7
1	15.7
2	13.6
3	11.5
4	10.3
5	8.9
6	8.0
7	7.3
8	6.6
9	6.1
10	5.7
11	5.3
12	4.9
13	4.5
14	4.3
15	4.1
16	3.9
17	3.7
18	3.5
19	3.4
20	3.2
21	3.0
22	2.9
23	2.7
24	2.6
25	2.4
26	2.3
27	2.1
28	2.0
29	1.8
30	1.6
31	1.4
32	1.3
33	1.1
34	1.0
35	0.8
36	0.7
37	0.6

After 38 minutes there was no acetone visible in the petri dish or under the carbon foam sample. The sample was placed in an air circulating oven at 60° C. to dry it and then allowed to cool to ambient temperature.

Example 15: Ethanol

Time (minutes)	Temperature (° C.)
0	21.6
1	20.3
2	19.6
3	19.0
4	18.6
5	18.1
6	17.8
7	17.4
8	17.1
9	16.9
10	16.7
11	16.5
12	16.3

-continued

Time (minutes)	Temperature (° C.)
13	16.2
14	16.0
15	15.8
16	15.7
17	15.6
18	15.5
19	15.4
20	15.3
21	15.1
22	15.0
23	15.1
24	14.9
25	14.8
26	14.8
27	14.8
28	14.7
29	14.7
30	14.6
31	14.6
32	14.6
33	14.5
34	14.5
35	14.4
36	14.4
37	14.4
38	14.3

After 38 minutes there was a significant amount of ethanol visible in the bottom of the petri dish. The sample was placed in an air circulating oven at 60° C. to dry it and then allowed to cool to ambient temperature.

Example 16: Water

Time (minutes)	Temperature (° C.)
0	20.9
1	20.3
2	20.2
3	20.1
4	19.9
5	19.8
6	19.7
7	19.6
8	19.5
9	19.5
10	19.5
11	19.5
12	19.5
13*	19.5
14	19.4
15	19.4
16	19.3
17	19.3
18	19.3
19	19.3
20	19.3
21	19.3
22	19.2
23	19.2
24	19.1
25	19.1
26	19.1
27	19.1
28	19.1
29	19.1
30	19.0
31	19.0
32	19.0
33	19.0
34	18.9
35	18.9

-continued

Time (minutes)	Temperature (° C.)
36	18.9
37	18.9
38	18.9

*Additional water squirted over carbon foam sample.

- 5 After 38 minutes there was a significant amount of water visible in the bottom of the petri dish.
 What is claimed is:
 1. A process comprising:
 (1) heating a liquified pitch under non-oxidizing, super-atmospheric conditions below about 500° C. sufficient to produce a carbon foam;
 15 (2) cooking said carbon foam by heating said carbon foam at a temperature above about 500° C. under non-oxidizing, superatmospheric conditions;
 20 (3) heating the coked carbon foam under conditions sufficient to produce an essentially graphitic carbon foam; and
 25 (4) at least partially filling porous regions in the essentially graphitic carbon foam with a phase change material.
 2. The process of claim 1 wherein the temperature in step (2) is held constant for at least 15 minutes.
 3. The process of claim 1 wherein the pressure is held substantially constant during each of steps (1) and (2).
 4. The process of claim 3 wherein the pressure in steps (1) and (2) is substantially identical.
 5. The process of claim 1 wherein the temperature during the cooking in step (2) is no more than about 1000° C.
 6. The process of claim 5 wherein at least some foaming in step (1) occurs at a temperature between about 420 and about 480° C. and wherein said pitch is selected from the group consisting of liquified mesophase petroleum pitch and liquified mesophase synthetic pitch.
 7. The process of claim 1 wherein said heating under conditions sufficient to produce an essentially graphitic carbon foam in step (3) includes a temperature in the range of about 2500° to 2800° C.
 8. The process of claim 1 wherein said heating in step (3) comprises heating the coked carbon foam from step (2) at a temperature above that for coking in step (2) to produce a carbonized carbon foam, followed by further heating at a yet higher temperature to produce said essentially graphitic carbon foam.
 9. The process of claim 1 wherein between steps (2) and (3) the cooked carbon foam is cooled.
 10. The process of claim 1 wherein said pitch is a liquified mesophase pitch.
 11. The process of claim 1 wherein said pitch is a liquified pitch of petroleum or synthetic origin.
 12. The process of claim 1 wherein said pitch is selected from the group consisting of liquefied mesophase petroleum pitch and liquified mesophase synthetic pitch.
 13. The process of claim 3 wherein said pitch is a liquified mesophase petroleum pitch.
 14. The process of claim 1 wherein at least some foaming in step (1) occurs at a temperature between about 4200 and about 480° C.
 15. The process of claim 14 wherein said pitch is a liquified mesophase pitch.
 16. The process of claim 14 wherein said pitch is a liquified pitch of petroleum or synthetic origin.
 17. The process of claim 1 wherein (a) at least some foaming in step (1) occurs at a temperature between about

420° and about 480° C. and (b) the temperature during the cooking in step (2) is no more than about 1000° C. and is held at a substantially constant value for at least 15 minutes.

18. The process of claim 17 wherein said pitch is selected from the group consisting of liquified mesophase petroleum pitch and liquified mesophase synthetic pitch.

19. The process of claim 17 wherein said pitch is a liquified mesophase pitch.

20. The process of claim 7 yielding in step (3) an essentially graphitic carbon foam having a bulk thermal conductivity between about 58 and 106 W/m² K.

21. The process of claim 7 yielding in step (3) an essentially graphitic carbon foam having a specific thermal conductivity greater than four times that of copper.

22. The process of claim 7 yielding in step (3) an essentially graphitic carbon foam having a specific thermal conductivity greater than copper.

23. The process of claim 7 yielding in step (3) an essentially graphitic carbon foam characterized by an X-ray diffraction pattern substantially as depicted in FIG. 14.

24. The process of claim 7 to yielding in step (3) an essentially graphitic carbon foam characterized by an X-ray diffraction pattern exhibiting doublet peaks at 2θ angles between 40 and 50 degrees.

25. The process of claim 7 yielding in step (3) an essentially graphitic carbon foam characterized by an X-ray diffraction pattern having an average d002 spacing of about 0.336.

26. The process of claim 25 wherein the average d002 spacing is 0.336.

27. The process of claim 20 wherein said pitch is a liquified mesophase pitch.

28. The process of claim 27 wherein at least some foaming in step (1) occurs at a temperature between about 420° and 480° C.

29. The process of claim 21 wherein said pitch is a liquified pitch of petroleum or synthetic origin.

30. The process of claim 29 wherein at least some foaming in step (1) occurs at a temperature between about 420° and 480° C. and the temperature during the cooking in step (2) is no more than about 1000° C.

31. The process of claim 20 wherein said pitch is selected from the group consisting of liquified mesophase petroleum pitch and liquified mesophase synthetic pitch.

32. The process of claim 31 wherein at least some foaming in step (1) occurs at a temperature between about 420 and 480 C. and the pressure is held substantially constant during each of steps (1) and (2).

33. The process of claim 21 wherein said pitch is a liquified mesophase petroleum pitch.

34. The process of claim 33 wherein (a) at least some foaming in step (1) occurs at a temperature between about 420° and 480° C., (b) the temperature during the cooking in step (2) is no more than about 1000° C. and is held at a substantially constant value for at least 15 minutes; and (c) the pressure is held to substantially the same constant value in steps (1) and (2).

35. The process of claim 20 wherein at least some foaming in step (1) occurs at a temperature between about 420° and 480° C.

36. The process of claim 1 wherein the phase change material freezes at a temperature above 0° C. but does not boil below room temperature.

37. The process of claim 1 wherein the phase change material changes phase from a solid to a liquid below 1800° C. but remains in the liquid state up to a temperature as high as 1800° C.

38. The process of claim 1 wherein the phase change material changes phase from a solid to a liquid below 1200° C. but remains in the liquid state up to a temperature as high as 1200° C.

39. The process of claim 1 wherein the phase change material is selected from the group consisting of water, acetic acid, germanium, germanium-silicon, and paraffin wax.

40. A process comprising at least partially filling the porous region of a thermally conductive, non-oxidatively stabilized, mesophase pitch-derived carbon foam with a phase change material.

41. The process of claim 40 wherein the carbon foam has a bulk thermal conductivity between about 58 and 106 W/m² K.

42. The process of claim 40 wherein the carbon foam has a specific thermal conductivity greater than four times that of copper.

43. The process of claim 40 wherein the carbon foam has a specific thermal conductivity greater than copper.

44. The process of claim 40 wherein the carbon foam has a bulk thermal conductivity less than or equal to 106 W/m² K.

45. The process of claim 40 wherein the carbon foam has a specific thermal conductivity of at least about 109 W-cm³/m²K.g.

46. The process of claim 40 wherein the carbon foam has a specific thermal conductivity of between about 109 and about 200 W-cm³/m²K.g.

47. The process of claim 40 wherein the carbon foam is characterized by an x-ray diffraction pattern substantially as depicted in FIG. 14.

48. The process of claim 40 wherein the carbon foam is characterized by an X-ray diffraction pattern having an average d002 spacing of about 0.336.

49. The process of claim 40 wherein the carbon foam is characterized by an X-ray diffraction pattern exhibiting doublet peaks at 2θ angles between 40 and 50 degrees.

50. The process of claim 49 wherein the average d002 spacing of the X-ray diffraction pattern is 0.336.

51. The process of claim 40 wherein the carbon foam, when initially at a temperature of approximately 21° C. and saturated with acetone, cools to a temperature of -50° C. in no more than about 4 minutes when under a reduced vacuum pressure of about 1.2 torr.

52. The process of claim 40 wherein the carbon foam, when initially at a temperature of approximately 21° C. and saturated with ethanol, cools to a temperature of -28° C. in no more than about 6 minutes when under a reduced vacuum pressure of about 1.0 torr.

53. The process of claim 40 wherein the carbon foam, when initially at a temperature of approximately 21° C. and saturated with water, cools to a temperature of -5.5° C. in no more than about 9 minutes when under a reduced vacuum pressure of about 29 torr.

54. The process of claim 40 wherein the carbon foam, when initially at a temperature of approximately 21° C. and saturated with acetone, cools to a temperature of 0.50° C. in no more than about 38 minutes when under ambient room temperature conditions.

55. The process of claim 40 wherein the carbon foam, when initially at a temperature of approximately 21° C. and saturated with ethanol, cools to a temperature of 14° C. in no more than about 38 minutes when under ambient room temperature conditions.

56. The process of claim 40 wherein the carbon foam, when initially at a temperature of approximately 21° C. and saturated with water, cools to a temperature of 19° C. in no more than about 38 minutes when under ambient room temperature conditions.